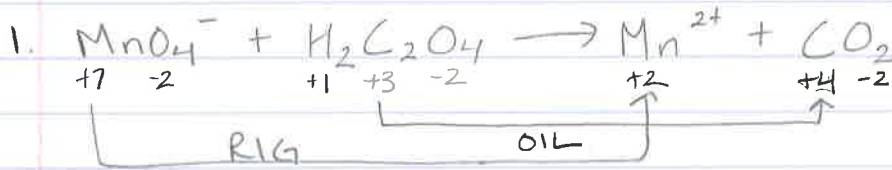


Electro Chem Practice



(sorry for the formatting issues!)

- 1 A
- 2 C
- 3 A
- 4 B
- 5 C
- 6 C



Able to oxidize all metals below it on chart



Won't oxidize metals ABOVE!

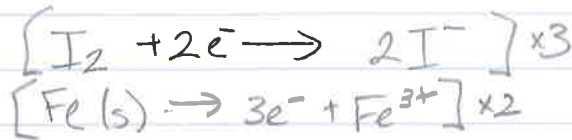
Ag(s), Cu(s) are between.

They will react with NO_3^- , but Not H^+

- 7 E
- 8 D
- 9 C

$$= 1.36 + 1.04$$

Fe^{3+} will oxidize Ni(s) $E = -0.036 + 0.23 \approx 0.2\text{V}$



} 6e⁻ to balance!

1) Ox half



(see above!)

RED half



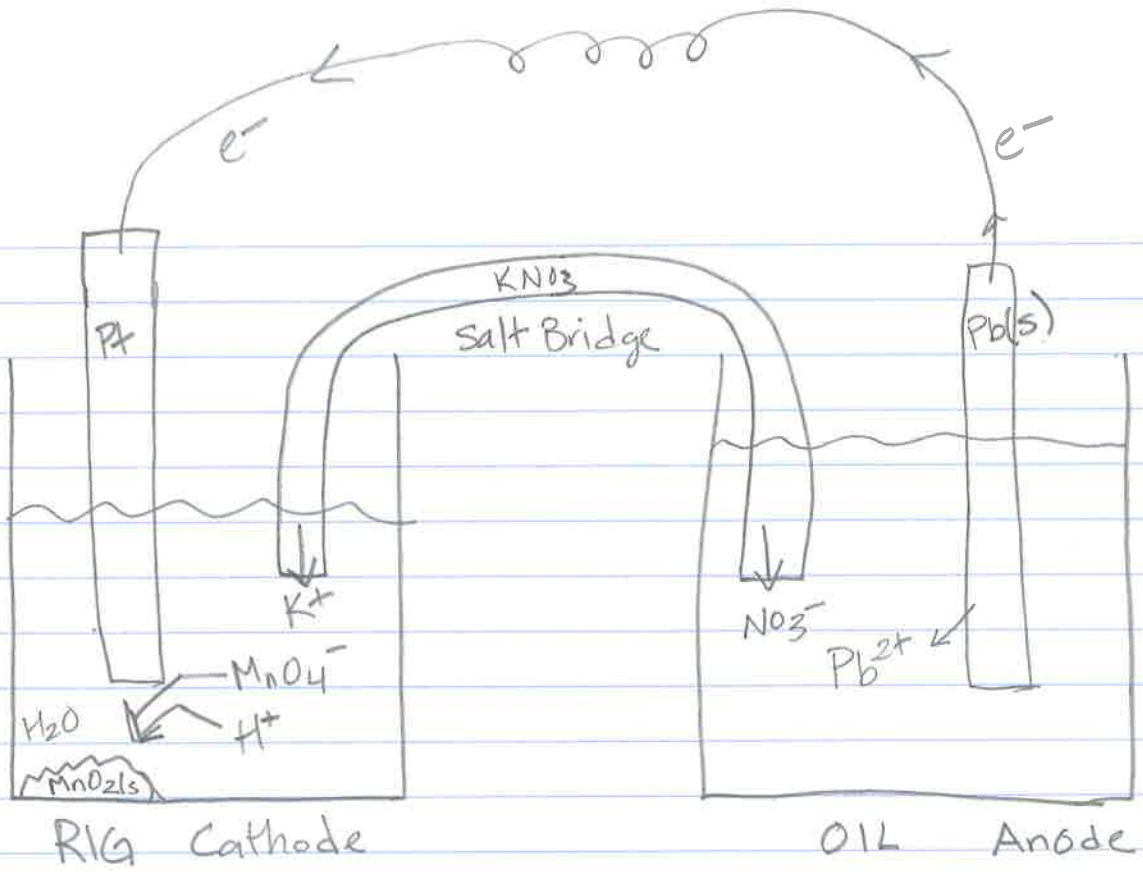
16-10



a) add OH^- to each side until All H^+ are neutralized.

FIVE STAR.
★★★★★

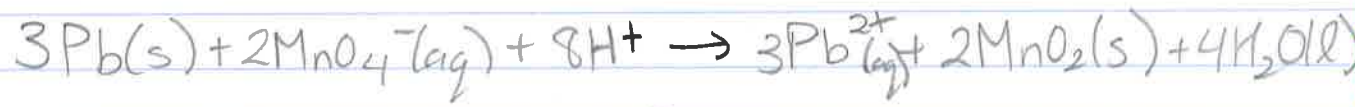
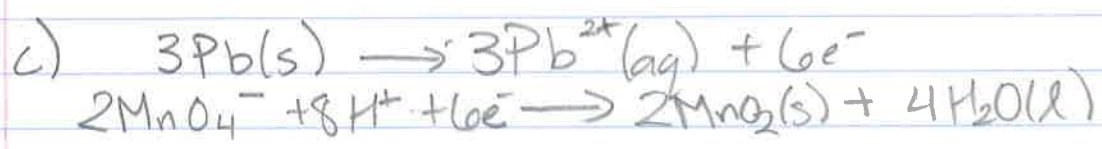
2)



FIVE STAR.
★★★★★

b) $E^\circ_{cell} = +.13V + 1.68V = \boxed{1.81V}$

FIVE STAR.
★★★★★



$Q = \frac{[Pb^{2+}]^3}{[H^+]^8 [MnO_4^-]^2} = 5.86 \times 10^{-6}$

FIVE STAR.
★★★★★

$E = E^\circ - \frac{.0592V}{6} \log Q = \boxed{1.86V}$

$\frac{1.81V}{- .0516}$