

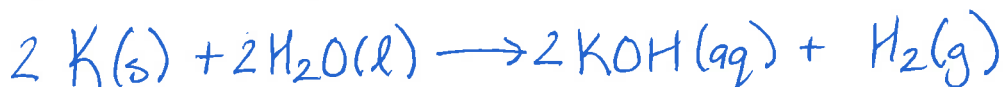
Ch 7 Reaction Practice Problems:

Write equations for and balance the following chemical reactions.

1. Aqueous nickel (II) chloride reacts with aqueous sodium hydroxide to produce nickel(II) hydroxide precipitate and aqueous sodium chloride.



2. Solid potassium metal reacts with water to give aqueous potassium hydroxide and hydrogen gas.



3. Aqueous sodium hydroxide reacts with aqueous phosphoric acid to give liquid water and aqueous sodium phosphate.

I Don't expect you to know this!
 H_3PO_4



4. Aqueous hydrogen fluoride reacts with aqueous potassium hydroxide to give water and aqueous potassium fluoride.



Complete and balance the following reactions, determining if a precipitate is formed. If both potential products are soluble, write no reaction.



Potential Products	Soluble or Insoluble
CsOH	insol
NaI	sol



Potential Products	Soluble or Insoluble
NaCl	sol
NH_4NO_3	sol



Potential Products	Soluble or Insoluble
KNO_3	sol
PbCl_2	insol



Potential Products	Soluble or Insoluble
AgBr	insol
$\text{Mg}(\text{NO}_3)_2$	sol



Potential Products	Soluble or Insoluble
CaCl_2	sol
$\text{Fe}(\text{OH})_3$	insol

Acid / Base



NOT
Precip!

