

Mixed Naming Key

1. CCl_4
2. $\text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2$
3. mercury (II) nitrate
4. sodium sulfite
5. $\text{Zn}(\text{OH})_2$
6. strontium nitrate
7. hydrogen monofluoride
8. Fe_2O_3
9. aluminum nitrate
10. calcium carbonate
11. SCl_2
12. $\text{Mg}_3(\text{PO}_4)_2$
13. barium sulfate
14. Li_2CrO_4
15. sodium cyanide
16. KClO_3
17. H_2O
18. potassium dichromate
19. sulfur trioxide
20. dihydrogen monosulfide
21. $\text{Ba}(\text{OH})_2$
22. $(\text{NH}_4)_3\text{PO}_4$
23. $\text{V}_2(\text{SO}_4)_5$
24. Li_2O
25. Sodium sulfide
26. Bismuth(III) fluoride
27. Copper (II) phosphide
28. Nitrogen monoxide
29. Li_3N
30. Hydrogen monobromide
31. $\text{Cr}_2(\text{CO}_3)_3$

32. $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$
33. N_2O_4 (There is a typo on the worksheet! Tetroxide, no a.)
34. Carbon dioxide
35. K_3PO_4
36. Lead(II) acetate
37. Hg_2Cl_2 (Hg_2^{2+} is a polyatomic cation called mercury(I). As long as it is not listed in your table, I don't expect you to memorize this. You should be able to use a polyatomic even if it is not on your list. I would have expected you to write HgCl for this answer if you did not know about Hg_2^{2+} .)
38. Tin(IV) acetate
39. MgCl_2
40. Calcium sulfate
41. Lithium chloride
42. Copper(II) phosphate
43. Barium chromate
44. Iron(II) bromide
45. Potassium oxide
46. Magnesium hydroxide
47. dinitrogen monoxide
48. dihydrogen monosulfide
49. copper(II) sulfate
50. magnesium sulfate
51. SF_2
52. SnF_2
53. NaH
54. Rubidium bromide
55. Cesium fluoride
56. Aluminum iodide
57. Nitrogen trifluoride
58. Dinitrogen tetrafluoride
- 59.
- 60.
- 61.

- 62.
- 63.
64. silicon tetrafluoride
65. ruthenium (III) nitrate
66. vanadium (V) oxide
67. $\text{Pd}_3(\text{SO}_4)_4$
68. IrN
69. Thallium (I) chloride
70. Nickel (III) phosphide
71. Iron(II) oxide
72. TiS
73. $\text{Ti}(\text{SO}_4)_2$
74. Chromium (II) oxide
75. Cesium sulfide
76. Rubidium phosphide
77. Ru_3PO_4
78. S_2F_{10}
79. Chlorine dioxide
80. BF_3