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CHEM 1301/ Homework 6

1. Classify each compound as Type 1 Ionic or Type 2 Ionic:

<i>Formula</i>	<i>Type?</i>	<i>Formula</i>	<i>Type?</i>
Ba(OH)₂		Li₂CO₃	
PbO		NaCl	
Pb(NO₂)₃		TiO₂	
Al₂O₃		NH₄NO₃	

2. Name the following Type 1 Ionic Compounds:

NaOH	
K ₃ N	
SrCl ₂	
KI	
InPO ₃	
Ga ₂ (CO ₃) ₃	

3. Write the formulas for the following Type 1 ionic compounds:

Sodium sulfide	
Strontium chloride	
Magnesium carbonate	
Aluminum nitrate	
Cesium bromide	
Calcium cyanide	

4. Name the following Type 2 Ionic Compounds:

TiO ₂	
MnCO ₃	
Fe(NO ₃) ₃	
Cu ₃ N ₂	
V ₂ S ₅	
CuCl	

5. Write the formulas for the following Type 2 ionic compounds:

Chromium (II) Nitrate	
Tungsten (III) Carbonate	
Iron(II) fluoride	
Lead(II) chromate	
Cobalt(IV) phosphate	
Bismuth(V) nitrite	

6. Classify each compound as Type 1 Ionic ,Type 2 Ionic, or Covalent:

<i>Formula</i>	<i>Type?</i>	<i>Formula</i>	<i>Type?</i>
Ca(OH) ₂		FeCO ₃	
N ₂ O		CO ₂	
ClO ₂		SnO ₂	
Al(NO ₃) ₃		NaF	

7. Name the following Covalent/Molecular Compounds:

N_2O_4	
SO_3	
CCl_4	
P_2O_5	
$ClBr$	
OF_2	

8. Write the formulas for the following covalent/molecular compounds.

selenium dichloride	
carbon tetrabromide	
dihydrogen sulfide	
dinitrogen pentoxide	
sulfur hexafluoride	
phosphorus trichloride	